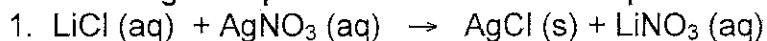
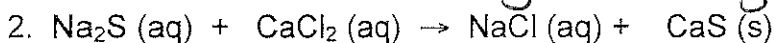
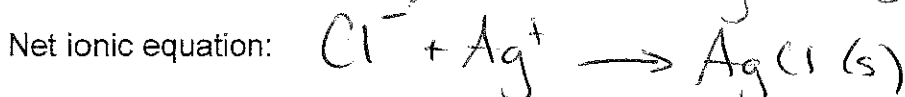
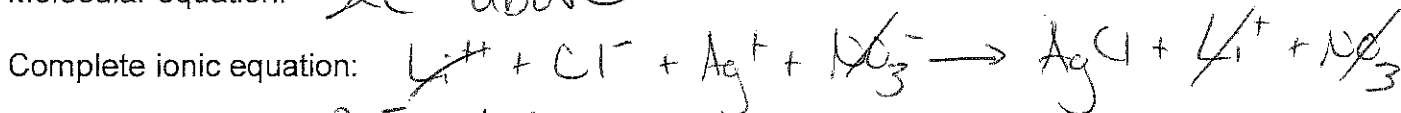


## Homework 3.1 – Practice Problems on Net Ionic Equations

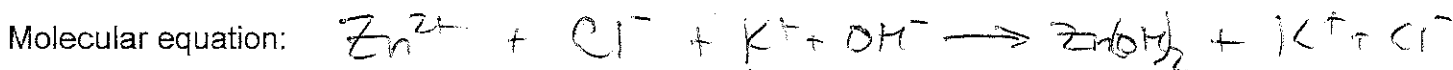
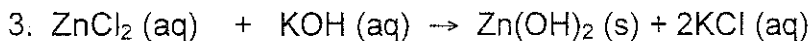
Show the total ionic and net ionic forms of the following equations. Note – You need to make sure the original equation is balanced before proceeding!



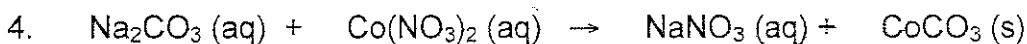
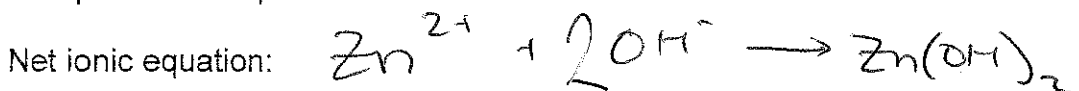
Molecular equation: *See above*



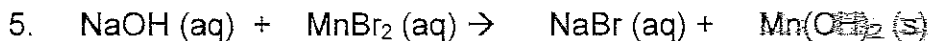
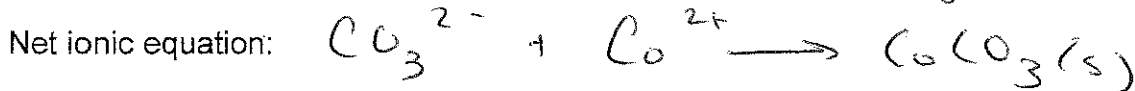
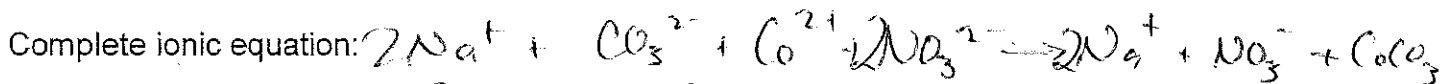
Molecular equation: *See above*



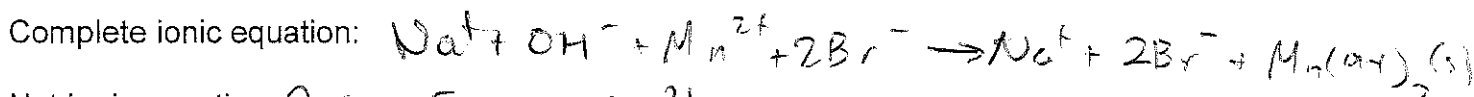
Complete ionic equation:



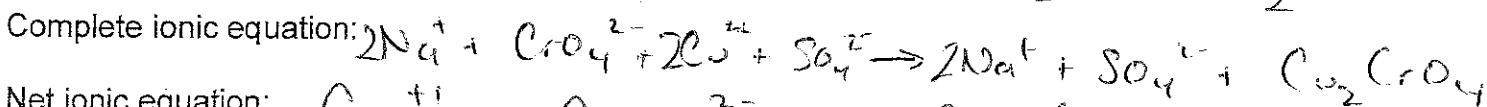
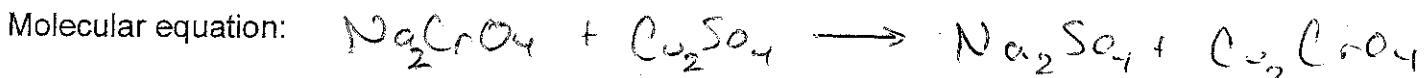
Molecular equation: *See above*



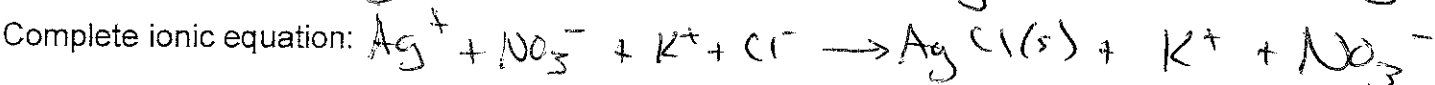
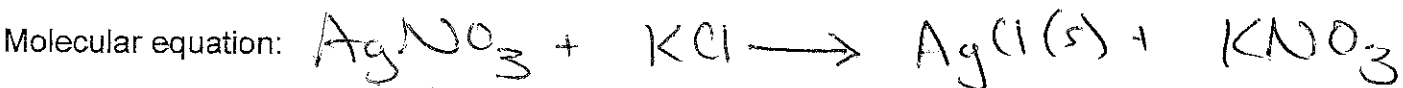
Molecular equation: *See above*



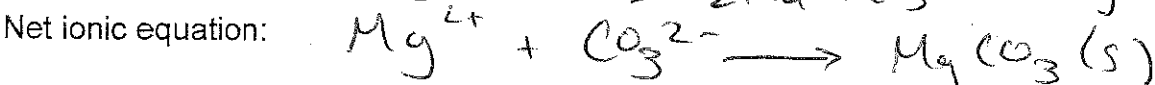
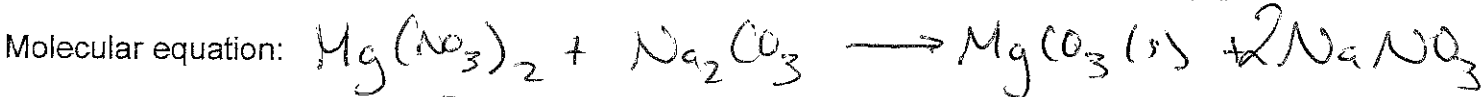
6. sodium chromate (aq) + Copper (I) sulfate (aq) → sodium sulfate (aq) + Copper Chromate (s)



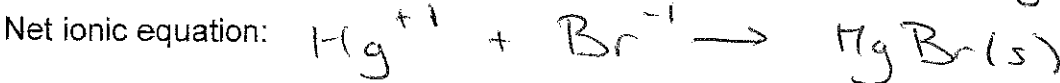
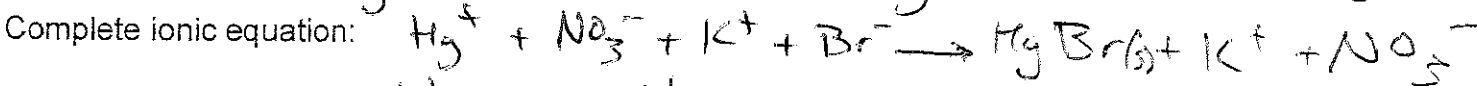
7. silver nitrate (aq) + potassium chloride (aq) → silver chloride (s) + potassium nitrate (aq)



8. magnesium nitrate (aq) + sodium carbonate (aq) → magnesium carbonate (s) + sodium nitrate (aq)



9. Mercury (I) Nitrate (aq) + Potassium Bromide (aq) → Mercury (I) Bromide (s) + Potassium Nitrate (aq)



10. Iron (III) Iodide (aq) + Silver Nitrate (aq) → Iron (III) Nitrate (aq) + Silver Iodide (s)

