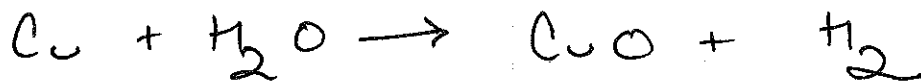


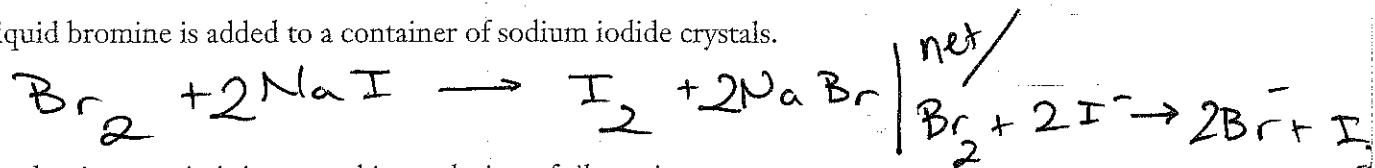
Exercise 8-1: Using the activity series, predict and balance the following single replacement reactions. Use abbreviations to indicate the appropriate phase of reactants and products where possible.

Note: Not all of the reactions will occur. For those that do not, write "no reaction."

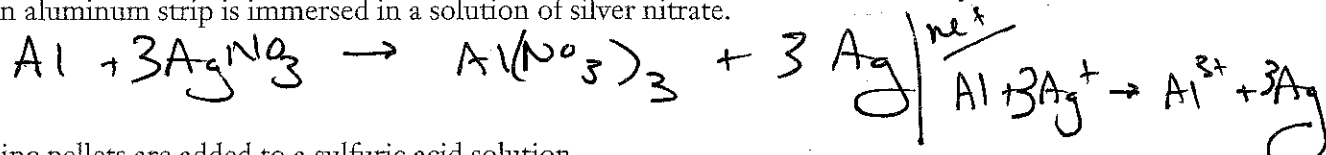
1. A piece of copper is dropped into a container of water.



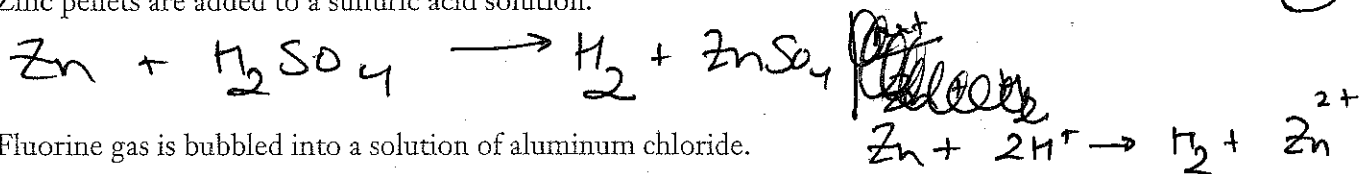
2. Liquid bromine is added to a container of sodium iodide crystals.



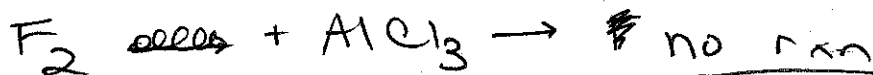
3. An aluminum strip is immersed in a solution of silver nitrate.



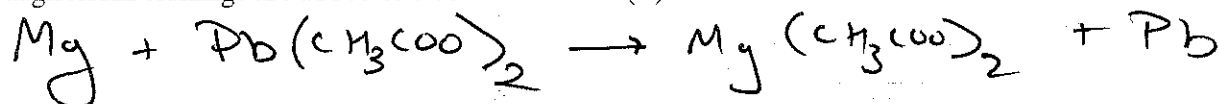
4. Zinc pellets are added to a sulfuric acid solution.



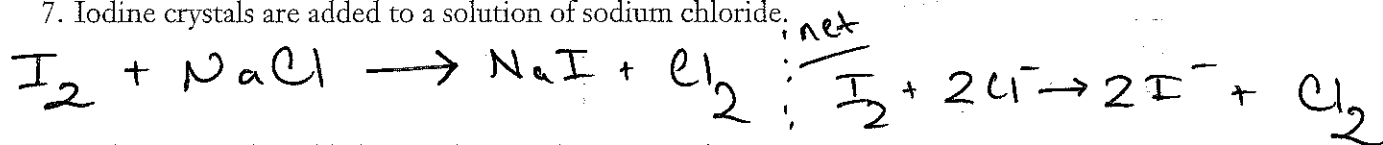
5. Fluorine gas is bubbled into a solution of aluminum chloride.



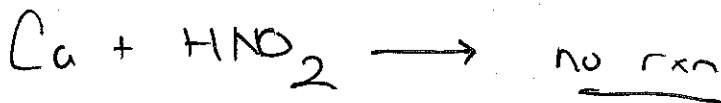
6. Magnesium turnings are added to a solution of lead(II) acetate.



7. Iodine crystals are added to a solution of sodium chloride.



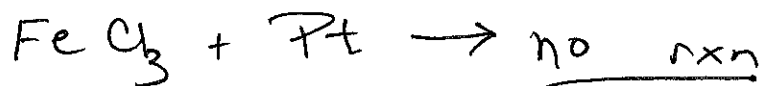
8. Calcium metal is added to a solution of nitrous acid.



9. A pea-sized piece of lithium is added to water.



10. A solution of iron(III) chloride is poured over a piece of platinum wire.



Note: On the AP reaction prediction section, all reactions "work"; in other words there will be no "No reactions" on the AP Exam.